

IN THE CLAIMS

Please amend the claims as follows:

1. (CURRENTLY AMENDED) A reversible hydrogen storage material comprising:

80 to 99.9 weight percent of an aluminum hydride; and

0.1 to 20 weight percent of a catalytic material comprising a reversible hydrogen storage alloy, a Raney catalytic material, or combinations thereof, adapted to increase said catalytic material increasing the kinetics of hydrogen absorption/desorption of said aluminum hydride without significantly reducing the hydrogen storage capacity of said aluminum hydride.

2. (CURRENTLY AMENDED) The reversible hydrogen storage material according to claim 1, wherein said aluminum hydride has the formula $X(AlH_4)$, wherein X is an element ~~chosen~~ selected from the group consisting of Group IA alkali metals, Group IIA alkali earth metals, Group IIIB lanthanides, ~~or~~ and Group IVB transition metals.

3. (ORIGINAL) The reversible hydrogen storage material according to claim 2, wherein X is Na, Li, Zr, or Mg.

4. (CANCELED)

5. (CURRENTLY AMENDED) The reversible hydrogen storage material according to claim 4, wherein said hydrogen storage alloy is selected from the group consisting of rare-earth/Misch metal alloys, zirconium alloys, titanium alloys, magnesium alloys, or combinations thereof.

6. (CURRENTLY AMENDED) The reversible hydrogen storage material according to claim 4, wherein said Raney catalytic material is selected from the group consisting of Raney nickel, Raney iron, Raney Cobalt, Raney Manganese, or combinations thereof.

7. (CURRENTLY AMENDED) A method of making a reversible hydrogen storage material comprising the steps of:

preparing a powder mixture comprising 80 to 99.9 weight percent of an aluminum hydride and 0.1 to 20 weight percent of a catalytic material comprising a reversible hydrogen storage alloy, a Raney catalytic material, or combinations thereof, ~~adapted to provide~~ said catalytic material providing said aluminum hydride with reversible hydrogen storage while not reducing the hydrogen storage capacity of said aluminum hydride; and

mechanically milling said mixture in an inert atmosphere.

8. (CURRENTLY AMENDED) The method according to claim 7, wherein said aluminum hydride has the formula $X(AlH_4)$, wherein X is

an element ~~chosen~~ selected from the group consisting of Group IA alkali metals, Group IIA alkali earth metals, Group IIIB lanthanides, or Group IVB transition metals.

9. (ORIGINAL) The method according to claim 8, wherein X is Na, Li, Zr, or Mg.

10. (CANCELED)

11. (CURRENTLY AMENDED) The method according to claim 10, wherein said hydrogen storage alloy is selected from the group consisting of rare-earth/Misch metal alloys, zirconium alloys, titanium alloys, magnesium alloys, or combinations thereof.

12. (CURRENTLY AMENDED) The method according to claim 10, wherein said Raney catalytic material is selected from the group consisting of Raney nickel, Raney iron, Raney Cobalt, Raney Manganese, or combinations thereof.